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Now we will balance the sodium atoms by placing a 6 in front of NaCl. Ca 3 N 2 + 6 NaCl 3 CaCl 2 + 2 Na 3 N 4. Finally, examine the chlorine atoms and notice that they are already balanced. 5. Double check each atom to make sure there are equal numbers of each on both sides of the equation. When balancing equations you NEVER change subscripts.

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24 ChemQuest 29 Monc LcW s Structures tr l Name Date: Hour: Introduction Questions Draw the Lewis structures for the following molecules. Part c is done for you. a) H2O (oxygen is central) b) SO2 c) N2 It t-t (a l)rt @ N=N @ @ A @ 2 In question 1 above you should see three different kinds of bonds as best you can. a) single bond: one es.iv (ee-) shave b) *-I; double bond:

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ChemQuest 29 Answer Key Hour l Introduction Questions Information: Subscripts and Coefficients A subscript is a small number that tells you how many atoms are in a compound.For example, in CaCl2 the two is the subscript and it tells us that there are two chloride ions bonded to one calcium. A coefficient tells also tells us how many atoms or compounds there are, but in a different way.

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7. Here are the answers to question 6. Double check your answers to question 6: A) CaCl2 + Li2O l CaO + 2 LiCl B) Al2S3 + 3 Cu l 3 CuS + 2 Al C) 2 Na3P + 3 MgBr2l Mg3P2 + 6 NaBr Information: Balancing Equations Containing Polyatomic Ions When an equation contains polyatomic ions, it may look a little more difficult to balance. But

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Balancing Equations: Answers to Practice Problems 1. Balanced equations. (Coefficients equal to one (1) do not need to be shown in your answers).

Balancing Equations: Practice Problems

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Chemquest 23 Answer Key More Lewis Structures Chemquest 23 Answer Key More Hint: add your answer to question 1 and four times your answer to question 2. 4 + 4(7) = 4 + 28 = 32 4. Verify that there are 32 electrons pictured in the drawing of CCl4.

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